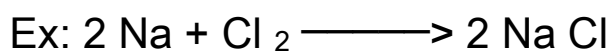
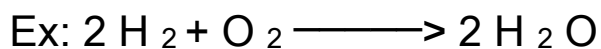
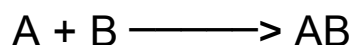


Types of chemical reaction

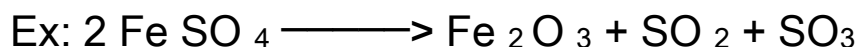
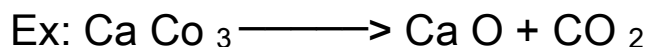
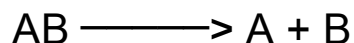
Some of the important types of chemical reactions are:

1. Combination reactions
2. Decomposition reactions
3. Displacement reactions
4. Double displacement reactions
5. Oxidation and Reduction reactions

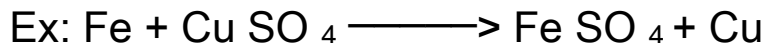
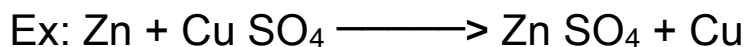
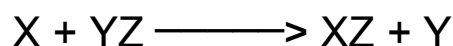
1. **Combination reactions:** A reaction in which two or more reactants combine to form a single product is known as a combination reaction. Combination reaction is also called synthesis reaction.



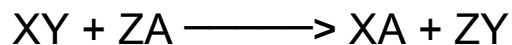
2. **Decomposition reactions:** Those reaction in which a compound splits up into two or more simpler substances are known as decomposition reactions.

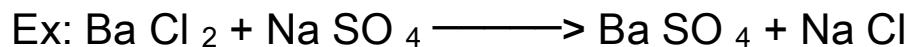


3. **Displacement reaction:** A chemical reaction in which a more reactive element displaces a less reactive element from its aqueous salt solution. It is also called a substitution reaction.



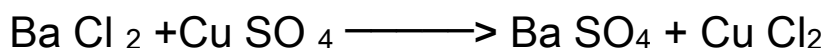
4. **Double Displacement Reaction:** A chemical reaction in which ions get exchanged between two reactants which form a new compound is called double displacement reaction. It is also called a metathesis reaction.





Precipitation Reaction: Any reaction in which an insoluble solid is formed that separates from the solution is called a precipitation reaction.

Ex: If barium chloride solution is added to copper sulphate solution, then a white precipitate of barium sulphate is produced alongwith copper chloride solution:



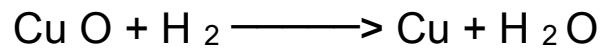
5. Oxidation and Reduction reactions:

Oxidation Reactions: (i) The addition of oxygen to a substance is called oxidation.

(ii) The removal of hydrogen from a substance is also called oxidation.

Reduction Reaction: (i) The addition of hydrogen to a substance is called reduction.

(ii) The removal of oxygen from a substance is also called reaction.



Redox Reaction: The oxidation and reduction reactions are also called redox reactions. In the name 'redox' the term 'red' stands for 'reduction' and 'ox' stands for oxidation.